

Visible Light Photocatalysis of [2+2] Styrene Cycloadditions by Energy Transfer**

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The prospect of conducting synthetically useful organic reactions with visible light has attracted significant recent attention from a number of research groups^[1] including our own.^[2] These efforts have focused on the utilization of the remarkable photoredox properties of ruthenium^[3] and iridium^[4] polypyridyl complexes, and a variety of photocatalytic transformations have been shown to occur upon irradiation with visible light in the presence of these complexes.^[5] The ability to use visible light rather than the ultraviolet light generally required for traditional organic photochemistry has numerous benefits,^[6] including: (1) the lower cost and decreased energy demand of visible light sources; (2) the ability to conduct photoreactions without specialized photo-reactors or quartz glassware; and (3) the ability to selectively photoexcite the transition metal photocatalyst without inducing undesired radical reactions of photochemically sensitive organic functional groups.

Our research group has been particularly interested in visible light photocatalysis of cycloaddition reactions.^[2] In our investigations, we have been able to exploit both photo-reduction and photooxidation reactions of Ru(bpy)₃²⁺ (**1a**) and related ruthenium(II) chromophores to design [2+2], [3+2], and [4+2] cycloaddition reactions involving radical anion and radical cation intermediates. Collectively, the diversity of products available using this strategy is quite broad; however, the nature of the photoinduced electron-transfer processes that generate the radical ion intermediates necessarily limits the scope of these reactions to either electron-deficient or electron-rich substrates that are amenable to one-electron redox processes. The success of photoredox methods reported from other labs has likewise been dictated by the redox properties of the organic substrates involved.

Recognizing that such electrochemical constraints will be important considerations in the design of any photoredox process, we wondered whether similar transformations could be initiated by *energy transfer* rather than by an electron-transfer mechanism. Although the quenching of Ru*(bpy)₃²⁺

by energy transfer has been documented with a number of organic compounds,^[7] to the best of our knowledge, there are only two carbon-carbon bond-forming reactions using transition metal photocatalysts that have been demonstrated to proceed via triplet sensitization of an organic substrate. The first is the Ru(bpy)₃²⁺-mediated norbornadiene-to-quadracyclane valence isomerization studied by Kutal for solar energy storage applications.^[8] The second is the photocatalytic dimerization of anthracene reported by Castellano to be sensitized by the related ruthenium photocatalyst Ru(dmb)₃²⁺ (dmb = 4,4'-dimethyl-2,2'-bipyridine).^[9] Thus, although the utility of UV-absorbing organic chromophores as triplet photosensitizers has been well established for decades,^[10] synthetic applications of triplet sensitization with transition metal complexes that absorb in the visible range have not been extensively explored.

We initiated our investigation by exploring the [2+2] photocycloaddition of styrene **3**, a substrate whose oxidation potential (+1.42 V vs SCE)^[11] has precluded its ability to participate in radical cation cycloadditions previously reported from our lab. Fluorinated iridium complex **2**, first reported by Malliaras and Bernhard^[12] and subsequently identified by Stephenson^[13] as an optimal visible light photocatalyst for Kharasch-type radical additions, does not possess an excited state oxidation potential (+0.89 V)^[12] sufficient to generate the radical cation of **1**. On the other hand, its reported emission maximum at 470 nm corresponds to an *E*_T of 61 kcal mol⁻¹. In general, styrenes possess excited state triplet energies (*E*_T) of approximately 60 kcal mol⁻¹.^[14] Together, these data suggested that **2** might be capable of sensitizing triplet-state reactions of styrene **3** upon irradiation with visible light.

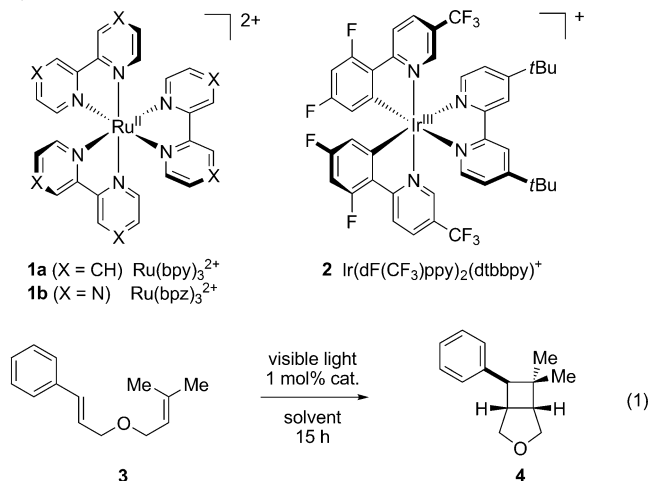
Indeed, irradiation of **3** with a 23 W compact fluorescent light bulb in the presence of 1 mol % of iridium complex **2**·PF₆ resulted in the formation of [2+2] cycloadduct **4** in a wide range of solvents (Table 1, entries 1–7). Consistent with a triplet sensitization mechanism, we observed no dramatic dependence on solvent polarity, whereas we have observed that radical cation processes benefit from the ability of polar solvents to stabilize the charged intermediates. Slightly faster conversion was observed in DMSO (entry 7), and upon optimization we were able to obtain the cycloadduct in good yield at lower reaction concentrations (entry 8). Also consistent with a triplet sensitization mechanism was the observation that metal complexes reported in other visible light photocatalysis applications with triplet state energies lower than that of styrene were ineffective in this transformation (entries 9 and 10).^[15] Finally, control reactions confirmed that no reaction occurs either in the absence of light or the absence of photocatalyst (entries 11 and 12).^[16]

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[**] This research was conducted using funds from the NIH (GM095666) and the Sloan Foundation. The NMR facilities at UW-Madison are funded by the NSF (CHE-9208463, CHE-9629688) and NIH (RR08389-01, RR13866-01).

Supporting information for this article is available on the WWW under <http://dx.doi.org/10.1002/ange.201204835>.

Table 1: Optimization and control studies for photocatalytic [2+2] cycloaddition.

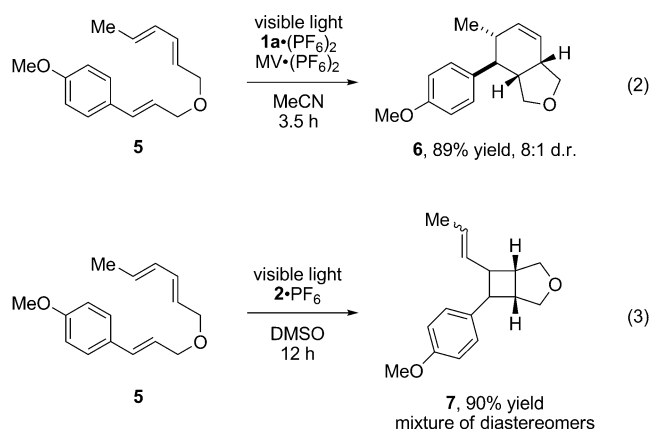


Entry ^[a]	Catalyst	Solvent	Conc. [M]	Yield ^[b] [%]
1	2·PF ₆	CHCl ₃	0.05	22
2	2·PF ₆	CH ₂ Cl ₂	0.05	26
3	2·PF ₆	THF	0.05	21
4	2·PF ₆	acetone	0.05	24
5	2·PF ₆	MeOH	0.05	21
6	2·PF ₆	MeCN	0.05	13
7	2·PF ₆	DMSO	0.05	33
8	2·PF ₆	DMSO	0.01	89 (83)^[c]
9	1a·(PF ₆) ₂	DMSO	0.01	0
10	1b·(PF ₆) ₂	DMSO	0.01	0
11	none	DMSO	0.01	0
12 ^[d]	2·PF ₆	DMSO	0.01	0

[a] Reactions irradiated using a 23 W compact fluorescent light bulb.
[b] Yields determined by ¹H NMR analysis against a calibrated internal standard unless noted. [c] Yield of the isolated product in parenthesis.
[d] Control reaction conducted in the dark.

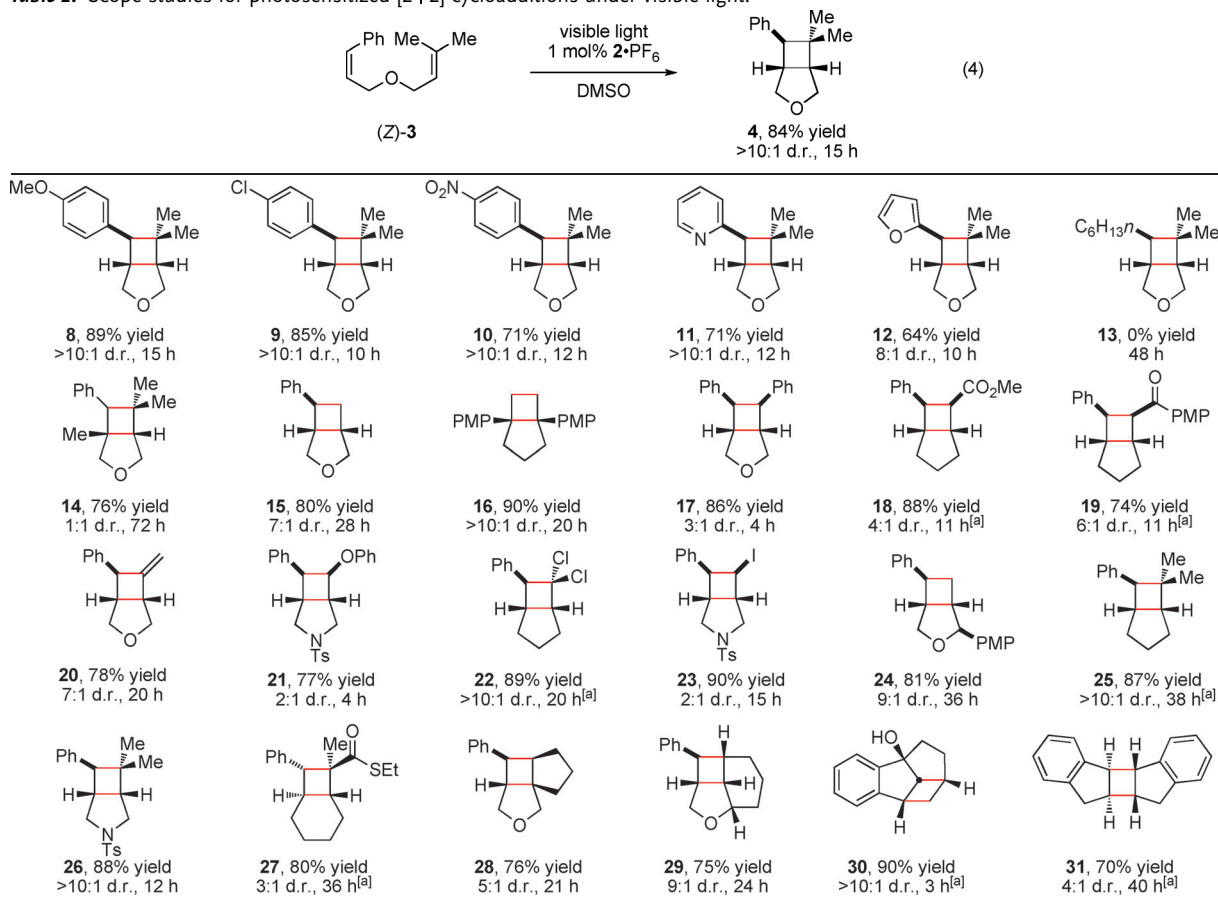
To further differentiate this triplet sensitization mechanism from radical cation cycloadditions, we investigated the reaction of diene substrate **5** under two different sets of conditions. First, Bauld has reported that the radical cation of **5** undergoes efficient intramolecular Diels–Alder cycloaddition,^[17] and we recently reported photocatalytic conditions that generate the same radical cation intermediate and produce [4+2] cycloadduct **6** in good yields [Eq. (2)].^[2h] On the other hand, exposure of **5** to iridium catalyst **2** under visible light affords a complex mixture of products that we identified as diastereomeric [2+2] cycloadducts [Eq. (3)],^[18] with only a trace of the [4+2] cycloadduct. Similar preferences for [2+2] periselectivity in cycloadditions involving dienes with a variety of triplet sensitizers have been previously reported.^[19]

Studies exploring the scope of this process are summarized in Table 2. Several themes emerge from an examination of these data. First, as expected from the propensity of alkene triplets to undergo geometric isomerization, the photocycloaddition is stereoconvergent; the reaction of the (*Z*) isomer of **3** gave results identical to the reaction of (*E*)-**3** [Table 2, Eq. (4)]. A time course of the reaction confirms that (*E*)/(*Z*) photoisomerization of **3** occurs faster than productive

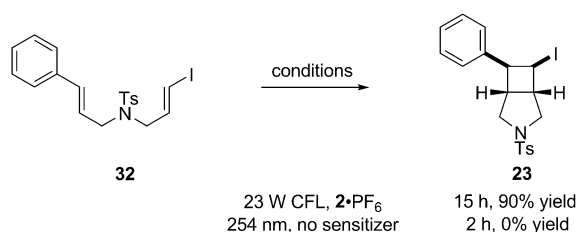


cycloaddition to **4** (see Supporting Information). Conversely, cyclic styrenes that are incapable of undergoing this energy-wasting isomerization react much more rapidly than acyclic styrenes (**30**). The ability of the cyclic constraint to extend the lifetime of the triplet is also presumably responsible for the ability of indene to undergo intermolecular cycloaddition (e.g., **31**); acyclic styrenes simply undergo rapid (*E*) to (*Z*) isomerization without productive intermolecular cycloaddition. We also observe that the reaction can be successfully initiated with a variety of substituted styrenes regardless of their electronic nature (**8–12**),^[20] including both electron-rich and electron-deficient heterostyrene compounds. On the other hand, substrates containing only aliphatic, non-conjugated alkenes are recovered unchanged from these reaction conditions (**13**), consistent with their significantly higher triplet energies compared to styrenes.^[21] The diversity of coupling partners suitable for this cycloaddition is extensive and includes aliphatic alkenes, styrenes, enones, enoates, enol ethers, haloalkenes, and allenes (**17–23**). The functional group compatibility of this process is also good; sulfonamides, unprotected alcohols, carbonyl compounds, and halogens are easily tolerated. Collectively, these data indicate that the range of cyclobutane compounds accessible using this methodology is quite broad and unconstrained by the redox properties of the substrates.

The ability to use low-energy, operationally convenient, and readily available visible light in the photosensitization reactions makes these processes more attractive for synthetic applications than UV irradiation. In order to demonstrate this point, we compared the reaction of vinyl iodide **32** under our optimized conditions to photoexcitation by direct irradiation with UV (Scheme 1). We find that **32** is rapidly consumed when irradiated in a Rayonet reactor outfitted with 254 nm lamps; however, the mass recovery of the reaction is poor, consistent with the propensity of high-energy UV radiation to promote uncontrolled radical decomposition processes. Indeed, resubjecting cyclobutane **23** to direct irradiation with 254 nm UV light resulted in its complete decomposition within 1 h. Thus, in addition to showing greater operational convenience than traditional UV photolyses, these conditions also increase the tolerance of photochemical reactions towards photosensitive functional groups such as alkyl iodides.

Table 2: Scope studies for photosensitized [2+2] cycloadditions under visible light.^[a]


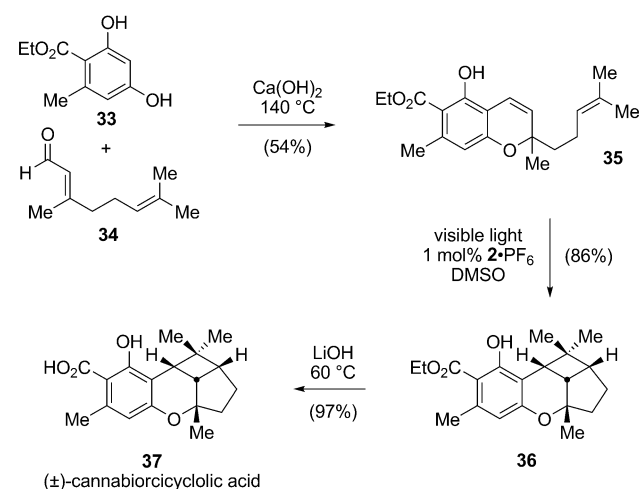
[a] Yields represent the averaged results of two reproducible experiments. Diastereomer ratios determined by ¹H NMR analysis of the unpurified reaction mixture. Reactions conducted using (E)-styrene isomer unless noted. PMP = *p*-methoxyphenyl. [b] Using (Z)-styrene isomer.


Scheme 1. Comparison of visible to UV conditions for cycloaddition of **32**. [a] ¹H NMR analysis against an internal standard revealed 100% consumption of **32**.

To demonstrate that these conditions are applicable to the preparation of complex natural products, we undertook a concise synthesis of cannabiorcycloic acid (**37**), one of several known cyclobutane-containing cannabinoids (Scheme 2).^[22] The chromene precursor **35** is easily obtainable by base-catalyzed condensation between phenol **33** and citral (**34**) in 54% yield. Photocycloaddition sensitized by 2-PF₆ affords 86% yield of cyclobutane **36** after 8 h of irradiation with a 23 W compact fluorescent light bulb. In contrast, the direct irradiation of **35** with 254 nm UV light for 5 h gave only 19% yield of the desired cycloadduct with only 9% of the starting chromene remaining. Finally, hydrolysis of the ethyl

ester under typical ester hydrolysis conditions affords (±)-cannabiorcycloic acid in 97% yield.

The results of this study are significant for a variety of reasons. First, they demonstrate that the class of transition metal photocatalysts that have earned significant recent interest for their ability to promote organic transformations


Scheme 2. Synthesis of cannabiorcycloic acid.

by photoinduced electron transfer can also catalyze reactions by energy transfer processes upon irradiation with household visible light sources. While this energy transfer cycloaddition maintains the same operational facility as previously reported photoredox transformations, the scope of the reaction is not constrained by the electrochemical properties of the substrates, and is instead governed by the relative excited state energies of the catalyst and styrenes. More broadly, these results suggest that a great variety of organic photochemical reactions known to be promoted by direct UV photoexcitation might also be accessible by visible light photocatalysis. Studies to investigate this proposition are currently underway in our laboratory.

Received: June 20, 2012

Published online: September 11, 2012

Keywords: cycloaddition · cyclobutanes · energy transfer · photocatalysis · visible light

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